



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS  
General Certificate of Education Ordinary Level

CANDIDATE  
NAME

CENTRE  
NUMBER

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CANDIDATE  
NUMBER

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**CHEMISTRY**

**5070/02**

Paper 2 Theory

**May/June 2008**

**1 hour 30 minutes**

Candidates answer on the Question Paper.

Additional Materials: Answer Booklet/Paper

**READ THESE INSTRUCTIONS FIRST**

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use a soft pencil for any diagrams, graphs or rough working.

Do **not** use staples, paper clips, highlighters, glue or correction fluid.

DO **NOT** WRITE IN ANY BARCODES.

**Section A**

Answer **all** questions.

Write your answers in the spaces provided in the Question Paper.

**Section B**

Answer any **three** questions.

Write your answers on any lined pages and/or separate answer paper.

A copy of the Periodic Table is printed on page 16.

The number of marks is given in brackets [ ] at the end of each question or part question.

At the end of the examination, fasten all your work securely together.

For Examiner's Use	
<b>Section A</b>	
<b>B7</b>	
<b>B8</b>	
<b>B9</b>	
<b>B10</b>	
<b>Total</b>	

This document consists of **16** printed pages.



**Section A**

Answer **all** the questions in this section in the spaces provided.

The total mark for this section is 45.

For  
Examiner's  
Use

**A1** Choose from the following gases to answer the questions below.

- ammonia**
- argon**
- carbon monoxide**
- chlorine**
- hydrogen**
- nitrogen**
- nitrogen dioxide**
- oxygen**

Each gas can be used once, more than once or not at all.

Name a gas which

**(a)** is made during the incomplete combustion of octane,  
..... [1]

**(b)** dissolves in water to make an alkaline solution,  
..... [1]

**(c)** is monatomic,  
..... [1]

**(d)** is a reducing agent in a Blast Furnace,  
..... [1]

**(e)** is used in the Contact process.  
..... [1]

[Total: 5]

**A2** Iron(II) sulphate,  $\text{FeSO}_4$ , is easily oxidised to iron(III) sulphate.

For  
Examiner's  
Use

(a) Calculate the percentage by mass of iron in iron(II) sulphate.

..... % [2]

(b) A sample of iron(II) sulphate is dissolved in water. Describe a test to show the presence of sulphate ions in this solution.

reagents .....

observation ..... [2]

(c) In the presence of aqueous hydrogen ions and dissolved oxygen, aqueous iron(II) ions are oxidised to form iron(III) ions and water.  
Write an ionic equation for this reaction.

..... [2]

(d) Aqueous iron(II) ions can also be oxidised by reaction with acidified potassium dichromate(VI),  $\text{K}_2\text{Cr}_2\text{O}_7$ . At the same time aqueous dichromate(VI) ions are reduced.

(i) Describe the colour change of the chromium-containing species during the reaction.

..... [1]

(ii) Describe the colour change of the iron-containing species during the reaction.

..... [1]

- (e) An impure sample of iron(II) sulphate was analysed by titration.

The sample was dissolved in  $25.0\text{ cm}^3$  of dilute sulphuric acid and then titrated against  $0.0400\text{ mol/dm}^3$  potassium dichromate(VI) solution.

$19.0\text{ cm}^3$  of potassium dichromate(VI) solution was required to reach the end-point.

- (i) Calculate the number of moles of potassium dichromate(VI) used in the titration.

..... moles [1]

- (ii) One mole of potassium dichromate(VI) reacts with six moles of iron(II) ions.  
Calculate the mass, in grams, of iron(II) ions in the sample analysed.

mass of iron(II) ions..... g [2]

[Total: 11]

For  
Examiner's  
Use

- A3** A student found a copy of a Periodic Table published in the year 1930. Several elements were missing from this table because they had not yet been discovered. One of these elements was technetium, Tc.

For  
Examiner's  
Use

One isotope of technetium has the symbol  ${}_{43}^{98}\text{Tc}$ .

- (a) Complete the table below to show the number of subatomic particles in one atom of this isotope.

number of protons	
number of electrons	
number of neutrons	

[2]

- (b) Suggest the symbol of another isotope of technetium.

..... [1]

- (c) Explain, in terms of subatomic particles and their charge, why an atom of  ${}_{43}^{98}\text{Tc}$  is electrically neutral.

.....  
 .....  
 ..... [2]

- (d) From its position in the modern Periodic Table predict two properties of technetium.

1 .....

2 ..... [2]

[Total: 7]

**A4** Ethane,  $C_2H_6$ , and ethene,  $C_2H_4$ , are both gaseous hydrocarbons.

For  
Examiner's  
Use

- (a) Describe how aqueous bromine can be used to distinguish between a sample of ethane and a sample of ethene.

.....  
 .....  
 ..... [2]

- (b) Draw a 'dot-and-cross' diagram for ethane.  
 You only need to draw the outer electrons of the carbon atoms.

[2]

- (c) Ethane reacts with chlorine in the presence of ultra-violet light.  
 Suggest a structure for a product of this reaction.

[1]

- (d) Write both the name and the molecular formula of an alkene molecule containing four carbon atoms.

name .....

molecular formula ..... [2]

[Total: 7]

- A5** One of the largest uses of phosphorus is in the making of safety matches. A safety match ignites when it is rubbed against the striking surface of a match box.

For  
Examiner's  
Use

The match head contains the following substances.

- phosphorus,  $P_4$
- potassium chlorate(V),  $KClO_3$
- sulphur, S
- a hydrocarbon wax

- (a)** The friction between the match head and the striking surface generates enough heat for the phosphorus to burn.  
Phosphorus burns to form phosphorus(V) oxide. This oxide is covalently bonded with a molecular structure.

- (i)** What is the molecular formula of phosphorus(V) oxide?

.....[1]

- (ii)** Suggest **one** physical and **one** chemical property of phosphorus(V) oxide.

physical property .....

.....

chemical property .....

.....[2]

- (b)** The heat from the combustion of phosphorus provides enough energy for the decomposition of potassium chlorate(V) to oxygen and potassium chloride.  
Construct the equation for the decomposition of potassium chlorate(V).

.....[2]

- (c)** The sulphur on the match head ignites.  
Write an equation to show the combustion of sulphur.

.....[1]

- (d)** Finally the wax on the match head begins to combust.  
One compound in the wax has the formula  $C_{18}H_{38}$ .  
To which class of hydrocarbons does this compound belong? Explain your answer.

.....

.....[1]

[Total: 7]

**A6** Sulphur dioxide,  $\text{SO}_2$ , and nitrogen dioxide,  $\text{NO}_2$ , are both atmospheric pollutants formed during the combustion of coal at a power station.

For  
Examiner's  
Use

**(a) (i)** State another source of sulphur dioxide as an atmospheric pollutant.

..... [1]

**(ii)** State another source of nitrogen dioxide as an atmospheric pollutant.

..... [1]

**(b)** Nitrogen dioxide and sulphur dioxide both cause acid rain. They are removed from the flue gases released from the power station by reaction with moist calcium carbonate in a process called flue gas desulphurisation.

Calcium carbonate reacts with sulphur dioxide to make a solid called calcium sulphite and a gas.

**(i)** What is the name of this gas?

..... [1]

**(ii)** Nitrogen dioxide reacts with calcium carbonate to make a solid. Suggest the name of this solid.

..... [1]

**(iii)** Describe one environmental effect of acid rain.

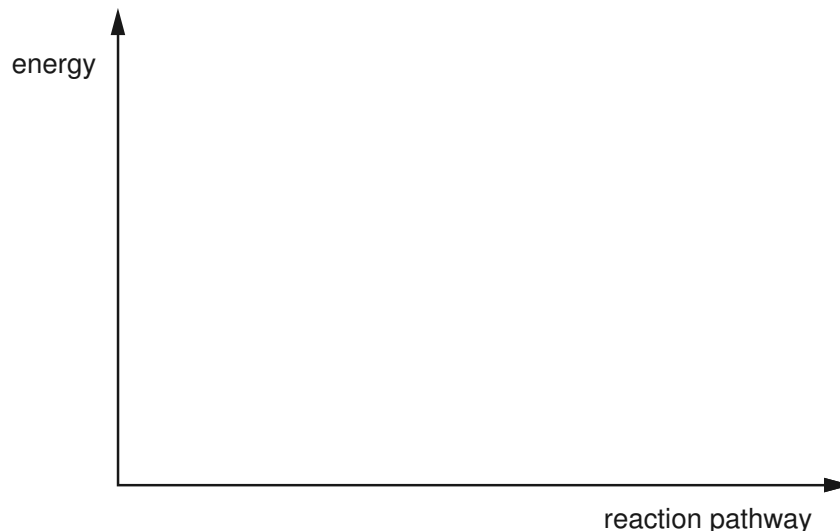
..... [1]

**(c)** Sulphur dioxide and nitrogen dioxide react together as shown in the equation.



Draw an energy profile diagram for this reaction.

Indicate both the enthalpy change and the activation energy on your diagram.



[3]

[Total: 8]



**Section B**

Answer **three** questions from this section.

The total mark for this section is 30.

For  
Examiner's  
Use

**B7** This question is about the chemistry of chlorine and some of its compounds.

- (a) Describe, with the aid of an ionic equation, the reaction of chlorine with aqueous potassium bromide. Explain why this reaction involves the reduction of chlorine. [3]
- (b) Magnesium reacts with chlorine to form magnesium chloride.  
Draw diagrams to show the electronic structures and charges of both ions present in magnesium chloride. [2]
- (c) Silver chloride is an insoluble salt.  
Outline the preparation of pure, dry silver chloride, starting from solid silver nitrate. [4]
- (d) State **one** environmental problem associated with the molecule  $C_2F_3Cl_3$ . [1]

[Total: 10]

- B8** Crude oil is a raw material which is processed in an oil refinery.  
Two of the processes used are fractional distillation and cracking.

For  
Examiner's  
Use

The table shows the percentage by mass of some different fractions in crude oil. The table also shows the demand for each fraction expressed as a percentage.

fraction	number of carbon atoms per molecule	percentage in crude oil	percentage needed by the oil refinery to supply demand
petroleum gases	1 - 4	4%	11%
gasoline	5 - 9	11%	22%
kerosene	10 - 14	12%	20%
gas oil	14 - 20	18%	15%
waxes and bitumen	over 20	23%	4%

- (a) The variation in which physical property is used to separate crude oil by fractional distillation? [1]
- (b) (i) Define the term *cracking*. [2]
- (ii) Use information from the table to explain how cracking helps an oil refinery match the supply of gasoline with the demand for gasoline. [2]
- (c) The hydrocarbon  $C_{15}H_{32}$  can be cracked to make propene and one other hydrocarbon.
- (i) Draw the structure of propene. [1]
- (ii) Write an equation for this reaction. [1]
- (d) Propene is used to make alcohols and poly(propene).
- (i) Describe how propene can be converted into an alcohol and draw the structure of this alcohol. [2]
- (ii) Draw the structure of poly(propene) showing at least two repeat units. [1]

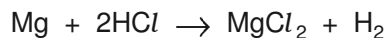
[Total: 10]

**B9** Dilute ethanoic acid and dilute hydrochloric acid both react with magnesium ribbon to form hydrogen.

For  
Examiner's  
Use

**(a)** Give the formula of one ion found in both of these dilute acids. [1]

**(b)** Magnesium ribbon reacts with hydrochloric acid as shown in the equation.



A 0.24 g sample of magnesium ribbon is added to 5.0 cm<sup>3</sup> of 2.0 mol/dm<sup>3</sup> hydrochloric acid.

**(i)** Which reactant, magnesium or hydrochloric acid, is in excess? Use calculations to explain your answer. [2]

**(ii)** Calculate the maximum mass of magnesium chloride that can be formed in this reaction. [2]

**(iii)** A 0.24 g sample of magnesium ribbon is added to 5.0 cm<sup>3</sup> of 2.0 mol/dm<sup>3</sup> ethanoic acid.

Explain why this reaction forms the same volume of hydrogen but takes place much more slowly than the reaction of the same mass of magnesium with 5.0 cm<sup>3</sup> of 2.0 mol/dm<sup>3</sup> hydrochloric acid. [3]

**(c) (i)** Write an equation for the reaction between dilute ethanoic acid and sodium carbonate. [1]

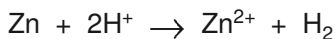
**(ii)** What observations would be made during this reaction? [1]

[Total: 10]

**B10** Brass is an alloy of zinc and copper.

For  
Examiner's  
Use

- (a) Describe, with the aid of a labelled diagram, the structure of a metal such as copper. [2]
- (b) Explain, in terms of their structures, why both zinc and copper are good conductors of electricity. [1]
- (c) A 1.2g sample of powdered brass was analysed by reaction with excess dilute sulphuric acid.  
The zinc reacts as shown in the equation to form 0.072dm<sup>3</sup> of hydrogen measured at room temperature and pressure.



- (i) Suggest why brass was used in a powdered rather than lump form. [1]
- (ii) Calculate the mass of zinc in the sample of brass. [2]
- (iii) Calculate the percentage of zinc in the sample of brass. [1]
- (d) Describe how aqueous ammonia can be used to show that only the zinc in the sample reacted with the acid. [3]

[Total: 10]







