



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
General Certificate of Education Ordinary Level

CANDIDATE
NAME

CENTRE
NUMBER

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CHEMISTRY

5070/21

Paper 2 Theory

October/November 2012

1 hour 30 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use a soft pencil for any diagrams, graphs or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

DO **NOT** WRITE IN ANY BARCODES.

Section A

Answer **all** questions.

Write your answers in the spaces provided in the Question Paper.

Section B

Answer any **three** questions.

Write your answers in the spaces provided in the Question Paper.

A copy of the Periodic Table is printed on page 20.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

For Examiner's Use	
Section A	
B7	
B8	
B9	
B10	
Total	

This document consists of **18** printed pages and **2** blank pages.



Section A

Answer **all** the questions in this section in the spaces provided.

For
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The total mark for this section is 45.

A1 (a) Define the term *element*.

..... [1]

(b) Choose from the following elements to answer the questions below.

aluminium

argon

bromine

gallium

helium

hydrogen

magnesium

nitrogen

oxygen

sodium

Each element can be used once, more than once or not at all.

Which element

(i) is in Group III and Period 4 of the Periodic Table, [1]

(ii) has atoms with 8 electrons in their outer shell, [1]

(iii) is a liquid at room temperature, [1]

(iv) reduces unsaturated vegetable oils to form a solid product, [1]

(v) forms an ionic chloride with the formula XCl_2 , [1]

(vi) is used in light bulbs? [1]

(c) Draw the electronic structure of an aluminium atom.

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[1]

[Total: 8]

A2 Steel is more resistant to corrosion than iron.

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(a) What are the essential conditions for the corrosion of iron?

.....
..... [1]

(b) Ships' hulls can be prevented from corroding by attaching pieces of magnesium to them. Explain why this prevents the hulls from corroding.

.....
..... [2]

(c) Steel is an alloy. Explain the meaning of the term *alloy*.

.....
..... [1]

(d) Samples of iron were placed in aqueous solutions having different pH values. The table shows how the speed of corrosion of iron varies with the pH of the solution.

speed of corrosion/cm per year	0.043	0.029	0.012	0.010	0.010	0.010	0.009	0.006
pH	2	3	4	5	6	8	10	12

Describe how pH affects the speed of corrosion of iron.

.....
.....
..... [2]

[Total: 6]

A3 The table below shows both the formulae and boiling points of the first five members of the alcohol homologous series.

For
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alcohol	formula	boiling point /°C
methanol	CH ₃ OH	65
ethanol	C ₂ H ₅ OH	79
propanol	C ₃ H ₇ OH	98
butanol	C ₄ H ₉ OH	117
pentanol	C ₅ H ₁₁ OH	138

(a) (i) Deduce the formula of the sixth member of the alcohol homologous series.

..... [1]

(ii) Predict the boiling point of this alcohol.

..... [1]

(b) Ethanol can be made industrially by fermentation.

Describe one other method of making ethanol industrially, stating the conditions required for the reaction.

.....

 [3]

(c) (i) Ethanol can be oxidised to ethanoic acid by atmospheric oxygen.
Name one other suitable oxidising agent which can be used.

..... [1]

(ii) Propanol can be oxidised to propanoic acid.
Draw the structure for propanoic acid.

[1]

[Total: 7]

A4 Water from natural sources, such as lakes and rivers, contains many dissolved substances.

(a) Name two dissolved substances that occur naturally in unpolluted water from lakes and rivers.

..... [1]

(b) Pollution in lakes and rivers can be caused by leaching of fertilisers from farmland. This can cause eutrophication.

(i) Name two ions present in fertilisers which cause eutrophication.

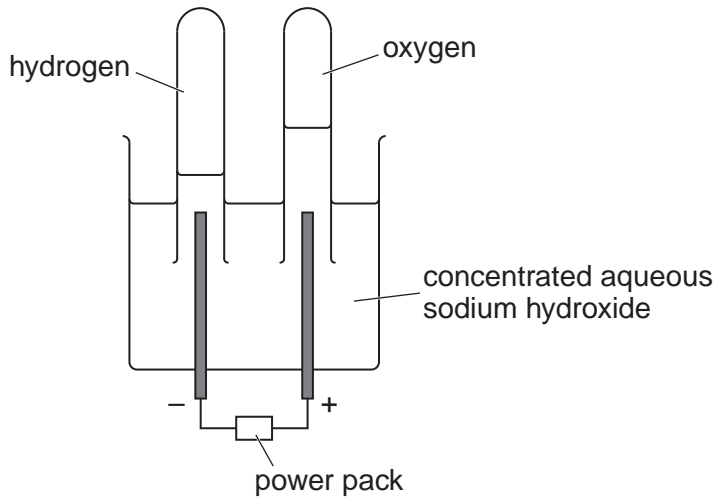
..... [2]

(ii) Describe the essential stages in eutrophication.

.....
.....
.....
.....
.....
.....
..... [4]

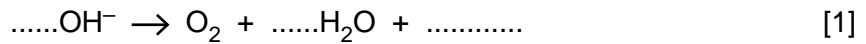
[Total: 7]

A5 The diagram below shows the apparatus used to electrolyse aqueous sodium hydroxide in the laboratory.



Electrolysis of the aqueous sodium hydroxide, results in the formation of hydrogen at the cathode (negative electrode) and oxygen at the anode (positive electrode).

(a) Complete the equation for the formation of oxygen at the anode.



(b) (i) When the power pack is replaced by a voltmeter, the apparatus acts like a fuel cell. The left hand electrode in the diagram becomes the negative pole of the cell and the right hand electrode becomes the positive pole.

State the direction of the electron flow in the external circuit.
Give a reason for your answer.

.....
..... [1]

(ii) In this fuel cell, hydrogen reacts with aqueous hydroxide ions to form water. Construct an equation for this reaction.

[1]

(c) (i) Suggest two advantages of using a fuel cell rather than petrol to power a car.

.....
.....
..... [2]

(ii) Suggest one disadvantage of fuel cells.

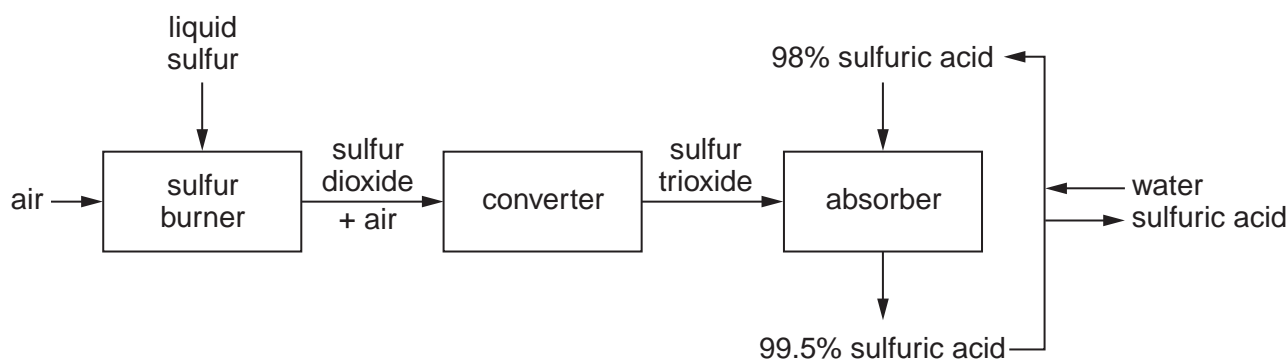
..... [1]

[Total: 6]

[Turn over

A6 A flow diagram for the manufacture of sulfuric acid is shown below.

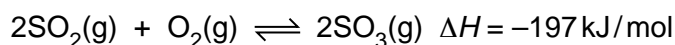
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- (a) In the sulfur burner, a spray of molten sulfur is burned in a furnace. Construct an equation for this reaction. Include state symbols.

[1]

- (b) In the converter, the following reaction occurs:



The yield of SO_3 is 95% at 450°C and atmospheric pressure.

- (i) Name the catalyst used in this reaction.

..... [1]

- (ii) Explain why increasing the pressure shifts the position of equilibrium further to the right.

..... [1]

- (iii) Explain why the reaction is carried out at atmospheric pressure even though an increase in pressure shifts the position of equilibrium further to the right.

..... [1]

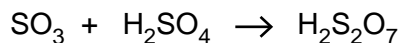
- (iv) Explain why the reaction is carried out at 450°C and not at a higher or lower temperature.

.....

 [3]

- (c) Sulfuric acid is formed from sulfur trioxide in two stages.
Firstly, the sulfur trioxide, SO_3 , is absorbed in concentrated sulfuric acid to form oleum, $\text{H}_2\text{S}_2\text{O}_7$.

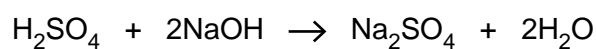
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The oleum is then mixed with water to form sulfuric acid.
Construct an equation for this reaction.

[1]

- (d) Aqueous sulfuric acid is titrated with aqueous sodium hydroxide.



It requires 28.0 cm^3 of 0.100 mol/dm^3 aqueous sodium hydroxide to neutralise 9.50 cm^3 of sulfuric acid.

Calculate the concentration, in mol/dm^3 , of the aqueous sulfuric acid.

Give your answer to 3 significant figures.

concentration of the aqueous sulfuric acid mol/dm^3 [3]

[Total: 11]

Section B

Answer **three** questions from this section in the spaces provided.

The total mark for this section is 30.

For
Examiner's
Use

B7 Tin is a metal in Group IV of the Periodic Table.

(a) Draw a labelled diagram to show the structure of a metal.

[2]

(b) Explain why metals

(i) conduct electricity,

(ii) are malleable.

..... [2]

(c) At high temperatures, tin reacts with steam to form tin(II) oxide, SnO, and one other product.

This reaction is reversible.

The other product is a gas which gives a 'pop' with a lighted splint.

(i) Construct an equation for this reaction.

[1]

(ii) Tin(II) oxide is an amphoteric oxide.
Explain the meaning of the term *amphoteric oxide*.

..... [1]

(d) (i) Concentrated nitric acid reacts with tin to form tin(IV) oxide, SnO₂, nitrogen dioxide and water.

Construct an equation for this reaction.

[1]

- (ii) Nitric acid contains nitrate ions.
Describe a test for nitrate ions.
Give the result of a positive test.

*For
Examiner's
Use*

.....
.....
..... [3]

[Total: 10]

B8 Petroleum is separated into fractions by fractional distillation.

For
Examiner's
Use

(a) Explain how fractional distillation separates petroleum into different fractions.

.....
.....
.....
.....
..... [3]

(b) The refinery gas fraction contains the first four members of the alkane homologous series.

(i) Explain the meaning of the term *homologous series*.

.....
.....
..... [2]

(ii) Draw the structure, showing all atoms and bonds, of the two isomers of butane, the fourth member of the alkane homologous series.

[2]

(c) Construct an equation for the complete combustion of hexane, C_6H_{14} .

[1]

(d) When long-chained alkanes are cracked in an oil refinery, shorter-chained alkanes and alkenes are formed.

For
Examiner's
Use

(i) Explain why the process of cracking needs to be carried out.

.....
..... [1]

(ii) Describe a chemical test to distinguish between an alkane and an alkene.

test
result [1]

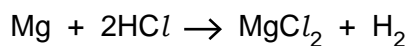
[Total: 10]

B9 (a) Define the term *relative atomic mass*.

.....
 [1]

For
Examiner's
Use

(b) The relative atomic mass of magnesium can be determined in the laboratory by finding the volume of hydrogen given off when magnesium reacts with hydrochloric acid.



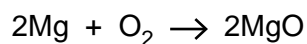
0.036 g of magnesium reacts at room temperature and pressure with excess hydrochloric acid to produce 36 cm³ of hydrogen.

1 mole of any gas at room temperature and pressure occupies 24 dm³.

Show by calculation that the relative atomic mass of magnesium is 24.

[3]

(c) Magnesium reacts with oxygen in the air to form magnesium oxide.



(i) If the yield of the reaction is 75% calculate the mass of magnesium oxide formed when 12 kg of magnesium burns in excess air.

[2]

(ii) Magnesium nitride is also formed when magnesium burns in air. Magnesium nitride is an ionic compound. Deduce the formula for magnesium nitride.

..... [1]

(d) When magnesium is heated with silicon, magnesium silicide, Mg_2Si , is formed. Magnesium silicide reacts with water to form silane, SiH_4 , and magnesium oxide.

For
Examiner's
Use

(i) Construct an equation for the reaction of magnesium silicide with water.

[1]

(ii) Silane has a structure similar to methane. Draw a 'dot-and-cross' diagram for silane. Show only the outer shell electrons.

[1]

(iii) Silane reacts with oxygen to form silicon dioxide and water. Construct an equation for this reaction.

[1]

[Total: 10]

B10 Limestone consists mainly of the compound calcium carbonate.

For
Examiner's
Use

- (a) Explain why limestone is used in the blast furnace for the extraction of iron. Include any relevant equations in your answer.

.....

.....

.....

.....

.....

..... [3]

- (b) Group II carbonates decompose on heating. The temperatures at which some Group II carbonates decompose are given in the table below.

Group II carbonate	decomposition temperature /°C
barium carbonate	1360
calcium carbonate	900
magnesium carbonate	540
strontium carbonate	1280

- (i) Which one of these carbonates is least likely to decompose on heating?
..... [1]
- (ii) Describe how the thermal stability of these carbonates changes with the reactivity of the metal.
..... [1]

(c) The speed of reaction of calcium carbonate with hydrochloric acid can be calculated by measuring the volume of gas given off at various time intervals.

For
Examiner's
Use

(i) Draw a labelled diagram of the apparatus you could use to follow the course of this reaction.

[2]

(ii) State and explain the effect of the following on the volume of a fixed mass of gas

- increasing the pressure,
- increasing the temperature.

.....

.....

.....

..... [3]

[Total: 10]

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DATA SHEET
The Periodic Table of the Elements

Group																	
I	II	III	IV	V	VI	VII	O										
1 H Hydrogen																	2 He Helium
3 Li Lithium	4 Be Beryllium	5 B Boron	6 C Carbon	7 N Nitrogen	8 O Oxygen	9 F Fluorine	10 Ne Neon										
11 Na Sodium	12 Mg Magnesium	13 Al Aluminium	14 Si Silicon	15 P Phosphorus	16 S Sulfur	17 Cl Chlorine	18 Ar Argon										
19 K Potassium	20 Ca Calcium	21 Sc Scandium	22 Ti Titanium	23 V Vanadium	24 Cr Chromium	25 Mn Manganese	26 Fe Iron	27 Co Cobalt	28 Ni Nickel	29 Cu Copper	30 Zn Zinc	31 Ga Gallium	32 Ge Germanium	33 As Arsenic	34 Se Selenium	35 Br Bromine	36 Kr Krypton
37 Rb Rubidium	38 Sr Strontium	39 Y Yttrium	40 Zr Zirconium	41 Nb Niobium	42 Mo Molybdenum	43 Tc Technetium	44 Ru Ruthenium	45 Rh Rhodium	46 Pd Palladium	47 Ag Silver	48 Cd Cadmium	49 In Indium	50 Sn Tin	51 Sb Antimony	52 Te Tellurium	53 I Iodine	54 Xe Xenon
55 Cs Caesium	56 Ba Barium	57 La Lanthanum	72 Hf Hafnium	73 Ta Tantalum	74 W Tungsten	75 Re Rhenium	76 Os Osmium	77 Ir Iridium	78 Pt Platinum	79 Au Gold	80 Hg Mercury	81 Tl Thallium	82 Pb Lead	83 Bi Bismuth	84 Po Polonium	85 At Astatine	86 Rn Radon
87 Fr Francium	88 Ra Radium	89 Ac Actinium															

140 Ce Cerium	141 Pr Praseodymium	142 Nd Neodymium	143 Pm Promethium	144 Nd Neodymium	145 Sm Samarium	146 Eu Europium	147 Pm Promethium	148 Gd Gadolinium	149 Tb Terbium	150 Dy Dysprosium	151 Ho Holmium	152 Er Erbium	153 Tm Thulium	154 Yb Ytterbium	155 Lu Lutetium
90 Th Thorium	91 Pa Protactinium	92 U Uranium	93 Np Neptunium	94 Pu Plutonium	95 Am Americium	96 Cm Curium	97 Bk Berkelium	98 Cf Californium	99 Es Einsteinium	100 Fm Fermium	101 Md Mendelevium	102 No Nobelium	103 Lr Lawrencium		

* 58–71 Lanthanoid series
† 90–103 Actinoid series

a = relative atomic mass
X = atomic symbol
b = atomic (proton) number

a	X	b
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Key

The volume of one mole of any gas is 24dm³ at room temperature and pressure (r.t.p.).